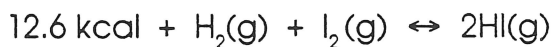
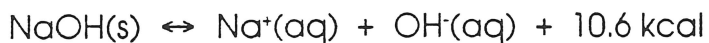


LE CHATELIER'S PRINCIPLE CONTINUED

Name _____



Stress	Equilibrium Shift	[H ₂]	[I ₂]	[HI]	K
1. Add H ₂	right	_____	decreases	increases	remains the same
2. Add I ₂			_____		
3. Add HI				_____	
4. Remove H ₂		_____			
5. Remove I ₂			_____		
6. Remove HI				_____	
7. Increase Temperature					
8. Decrease Temperature					
9. Increase Pressure					
10. Decrease Pressure					



(Remember that pure solids and liquids do not affect equilibrium values.)

Stress	Equilibrium Shift	Amount NaOH(s)	[Na ⁺]	[OH ⁻]	K
1. Add NaOH(s)		_____			
2. Add NaCl (Adds Na ⁺)			_____		
3. Add KOH (Adds OH ⁻)				_____	
4. Add H ⁺ (Removes OH ⁻)				_____	
5. Increase Temperature					
6. Decrease Temperature					
7. Increase Pressure					
8. Decrease Pressure					